

Martingrove Night School

Course : SCH 3U – Grade 11 (University) Chemistry

Instructor: R. Ng

Email: mossy_ng@ yahoo.ca



Textbook: online notes by Ms. Ng or McGraw Chemistry 11

<http://www.myteacherpages.com/webpages/RNg1>

Topics covered:

1. Nomenclature (naming)
 2. Stoichiometry (ratios)
 3. Gas laws
 4. Organic chemistry (the chemistry of Carbon)
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Marking scheme:

Term Mark (Quiz, Test, Assignments) = 70 %

Exam = 30 %

All evaluations will be divided equally (ie: 25%) per category

KU = COMM = APP = TIPS = 25 %

All marks are posted on Ms. Ng's website under My Links


Each student will have their own homepage and assigned their own password and ID to access their marks 24-7.

Instructions to access your marks on the web:

Click on My links on www.myteacherpages.com/webpages/RNg1

1. Alternatively, go to www.eclassinfo.com/home.asp?id=RSYNg
you will see

Ruth Ng at George Harvey CI
Listed below are the web resources that have
been published by your instructor.

 **WebGrade Reports**
[SCH4C](#)
[SCH4U\(Sem2\)](#)
[sch3u\(sem2\)](#)
[SCH 3U \(night\)](#)

2. Click on the tab, “ SCH 3U (Night) ”
You will see :

SCH 3U (night)

Use the ID and Password
provided by your instructor

Student ID

Password

3. Enter your ID (last name) and password (1st name) and LOGIN

*** You need to change your password and ID by emailing me.

You will see :

SCH 3U (night) - Sem 2
Instructor: Ruth Ng Student: Nguyen, Beatrice
Welcome to Ms. Ng Gr. 11 Chemistry class Click here to send email to instructor.
<small>[Click on a report to view it. You may need to disable pop-up blockers for this web site.]</small>
Student Summary Report: Student Summary
Regular Reports: Class Roster Class Standards

4. Click on “Student summary” to see your marks.

Lesson Sequence for SCH 3U

Lesson #	Topic covered
1	Naming of elements, ionic compounds (mono and polyatomic species)
2	Q1: Elements and polyatomic naming Naming of molecular species, acids Solubility of compounds
3	Q2: Solubility + acid/ molecular naming Balancing equations and identification of reaction types The Mole and Avogadro's number Calculating molecular weight Finding # of moles of a substance
4	Q3: Calculating moles + balancing equations Drawing shapes of molecules (Lewis and VSEPR) Intermolecular forces between molecules Types of bonding (ionic, molecular, metallic)
5	Q4: intermolecular forces + shapes Writing Net ionic equations (NIE) (identification of solid, (g), acids and complex ions)
6	Q5: Writing NIE Stoichiometry (basic ratios) Limiting reagent
7	Q6: Stoic and Limiting reagents % yield theory of the atom (S, p, d, f)

Lesson #	Topic covered
8	Q7: % yield + theory of atom Electronic configuration... (transition metal)
9	Review for BIG test from L1 to 8
10, 11	Acid/ base theories Strong versus weak acids Dissociation equations Conjugate acid/base pairs Lewis acids Calculate pH with limiting reagent
12	Q8: Acids Organic (single, double, triple bonds) And their substituents
13, 14	Q9: alkanes, alkenes, alkynes Alcohol, ethers, aldehyde, ketones, carboxylic acids, esters
15	Q10: alcohols etc... Review for BIG test L10 to 14
16	BIG test L 10 to 14 (acids + organic) Gas laws (units for P, T, V, R)
17	Q11: Gas units Gas laws ($PV = nRT$) and calculations
18	Q12: $PV = nRT$ calculations Hess Law ($Q = mc T$)
19	Review L 1 to 8 (Nomenclature, Stoic)

Lesson #	Topic covered
20	Review L 9 – 14 (Organic review)
21	Review L 15 – 18 (Gas Laws)
22	Practice Exam # 1 (GHCI Jan 2005 Exam.. taken up in class)
23	Mini Exam # 1 (covers ½ of the course)
24	Practice Exam # 2 (GHCI Jun 2005 Exam.. taken up in class)
25	Mini Exam # 2 (covers ½ of the course)
26	Prelim Exam (GHCI Jan 2006 Exam)
27	Final review for exam